Cumulative Part (100 pts)

1. (10 pts) Draw a line structure diagram for acetate: $\text{CH}_3\text{CO}_2(-)$. Your structure must include all lone pairs and formal charges. C and H atoms may be shown either implicitly or explicitly for full credit.

Answer for 1:

2. (10 pts) Draw the structure of $(Z)$-1-chloro-2-pentene.

Answer for 2:

3. (10 pts) Draw the best structure of the E2 product of $(1R, 2S)$-1-bromo-2-methylcyclohexane and tBuOK.

Answer for 3:

4. (10 pts) Circle the chiral molecules; draw an X over the achiral molecules.

5. (10 pts) How many possible stereoisomers (enantiomers + diastereomers) does 4-methyl-3,5-heptanediol have? _______

How many possible stereoisomers (enantiomers + diastereomers) does 2-methyl-3,4-pentanediol have? _______
6. (10 pts) Upon treatment with H$_3$O(+) the bicyclic carbon framework at left rearranges to make a bicyclic alcohol. **Please draw the structure.**

7. (10 pts) Draw the high-energy conformation of this molecule.

8. (10 pts) Draw the S$_\text{N}$1 products of this optically pure starting material.


10. (10 pts) How many $^{13}$C signals are there in the following molecules?
11. (10 pts) Predict the product.

\[
\text{hv} \quad \text{solvent} \quad \text{Br}^+ + \text{N} \text{O} \text{O} \quad \rightarrow \quad \text{product}
\]

12. (10 pts) Predict the Diels-Alder product. Stereochemistry is very important!

\[
\text{heat} \quad \text{solvent} \quad \text{Endo} \quad \text{stereocenter} \quad \text{count carbons!}
\]

13. (9 pts) Draw three lowest-energy molecular orbitals of cyclopentadienyl anion. Label these X1-X3 with one being the lowest-energy orbital. Hint, two of them have the same energy, in lecture we counted bonding, antibonding and non-bonding components in molecular orbitals. You may look at the pi-system from above and draw the atomic p orbitals as circles.

Orbitals for 13.

(5 pts) Are all of three of these MO’s occupied? \[\text{\underline{\text{yes}}}\]

(6 pts) How many more pi-molecular orbitals are there? \[2\]
14. (10 pts) Draw a circle around the aromatic (unusually stable) molecules. Put an X through the unstable molecules.

![Aromatic and Unstable Molecules](image)

15. (10 pts) Think about the transition states of these two reactions. One readily happens the other does not. Explain why one occurs and the other does not in terms of aromaticity? For the brevity refer to the reactions as 15a and 15b.

15a: $\text{N}=\text{C}+\text{O} \rightarrow \text{N}=\text{C}+\text{O}$

15b: $\text{N}=\text{C}+\text{O} \rightarrow \text{C}=\text{N}$

15b has 6e- 4A1+2 rule aromatic transition state.
15a has a 4e- ++ state.

16. (5 pts) a. Complete the reaction energy diagram by using it to describe a reaction that has one intermediate. (6 pts) b. Use the ‡ sign to label the transition state(s). c. (4 pts) Would you expect that the reaction vessel gets warmer, colder or maintains the same temperature as the reaction occurs? Warmer
17. (10 pts) Use combinations of the orbitals of ethylene to construct the four molecular orbitals of butadiene. Label the orbitals X1-X4 with X4 being the highest energy MO. Again just look down on the pi-system from above and draw the orbitals as circles.

18. (10 pts) Explain the following. In the hydrogenation of cyclooctatetraene dienes and trienes are isolated if the reaction is stop before completion. However only benzene and cyclohexane are observed in the incomplete hydrogenation of benzene. Why?

benzene is less reactive (more stable) than the diene or alkene hydrogenation products.

18. (5 pts) How many pi-electrons are there in the following molecule? _____
Work Space! This page must stay with your exam! You may refer the grader to this page for partial credit.
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